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From monolayer to multilayer:
perylene-3,4,9,10-tetracarboxylic diimide derivative acting either
as growth inhibitor or growth enhancer

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perylene-3,4,9,10-tetracarboxylic diimide

Abstract

Fine control of electrografting kinetics of diazonium salts is of paramount importance, particularly when considering application of diazoniums for fabrication of 2D nanomaterials. In this work, we develop on controlled grafting of perylenediimide (PDI) moiety separated with a 12-carbon aliphatic chain from aryldiazonium. The particular design of the diazonium cation synthesized for this study allows fine-tuning of the surface coverage by simple adjustment of the applied potential. Indeed, according to the potential imposed at the working electrode, the PDI moiety can either enhance the charge propagation within the growing layer or consume the diazonium salt in the bulk solution *via* redox cross-reaction. With this approach, the surface functionalization can be restricted to a monolayer or a multilayer in a robust and elegant manner, obeying Langmuir or first-order kinetics of electrografting, respectively. The experimental observations are supported with in-situ spectroelectrochemical investigations aimed to differentiate reduction of PDI moieties in the deposited layer and the bulk solution. Tentative mechanistic scheme is proposed and numerical simulations are undertaken to rationalize the data.

1. Introduction

Grafting of diazonium salts is widely used for the tailored modification of conductive interfaces properties as it allows a covalent attachment of a wide range of electrochemically active moieties to any type of substrate.^{1,2} This has already been proven to be a very powerful approach in a wide range of practical applications such as corrosion,³ sensing,⁴ electrocatalysis,⁵ energy storage⁶ and others.^{1,7} Up to the present day, the main drawback of diazonium grafting still remains a poor control of the deposited layer thickness due to the high reactivity of generated aryl radical species.^{1,2} This puts serious constraints on the fabrication of 2D-nanomaterials *via* electrochemical grafting where a fine control of deposition kinetics is required.

A few strategies have been proposed in an effort to overcome this issue. For example, it was shown that hindering 3,5-positions of the diazonium with chemically inert tert-butyl^{8,9} groups prevents anchoring of aryl radicals to already grafted aryl groups, significantly inhibiting grafting of multilayers. Some authors also used calix[4]arene¹⁰ or 3D Ru(II)¹¹ based moieties, which protected the grafted layers from the attack of aryl radicals and additionally led to the self-organization of the layer structure. Another group of methods was based on the degradation of preformed multilayers rather than on controlling the grafting kinetics *via* diazonium modifications. In this way, monolayers of thiophenolate¹² and benzaldehyde¹³ were obtained from grafting of diazoniums with cleavable disulfide and hydrazine groups. Although very successful, mentioned approaches cannot be translated to grafting of any diazonium compound and are hardly compatible with the introduction of any functional group since the use of specific organic moieties is required.

Against this background, a more universal approach has been suggested based on the addition of a cross-redox inhibitor to a reactive mixture of any diazonium.¹⁴⁻¹⁷ The formation of 4-nitrobenzene (NB) layers was significantly limited in the presence of 2,2-

diphenylpicrylhydrazyl (DPPH) redox mediator at -0.5 V vs Ag/AgNO₃ (all further potentials are quoted vs Ag/AgNO₃).^{14,15} Chronoamperometric grafting demonstrated that the surface coverage of NB layers drops drastically to a monolayer after a threshold potential of 0 V, which was coherent with the onset of DPPH reduction.^{15,17} Further chronoamperometric studies on the NB layer formation in aerated acetonitrile (ACN) solvent without DPPH showed that the decrease of NB surface coverage was equally observed after ca. -0.6 V.¹⁶ In this case, the potential threshold coincided with the onset of oxygen reduction. The suggested mechanism of inhibition of multilayer formation included the reduction of diazonium salt in the bulk solution with the reduced form of the redox mediator (DPPH⁻ or O₂⁻).¹⁷

Herein we take a further step towards developing new approaches for a fine control of the grafting kinetics *via* incorporation of a redox mediator (perylene-diimide - PDI derivative) directly into the diazonium molecule (Fig. 1). Moreover, the presence of a redox group in a grafted layer can render the deposited layer conductive and favor formation of thick multilayers as shown with 4-nitrobenzenediazonium, 4-benzoylbenzenediazonium and 9,10-dioxo-9,10-dihydroanthracene-2-diazonium.¹⁸ Therefore, we anticipate the opposite effects generated by a PDI group, resulting in either inhibition *via* a cross-redox reaction or acceleration of multilayer formation *via* a facilitation of radical formation due to enhanced heterogeneous electron transfer. Given the above background, the primary goal of this work was to investigate the kinetics of diazonium layer formation composed of a redox group, using PDI moiety as a typical system. The PDI and aryl groups were separated with an inert aliphatic chain (12 atoms) to suppress their electronic interaction and thus, preserve the reduction properties of individual moieties.

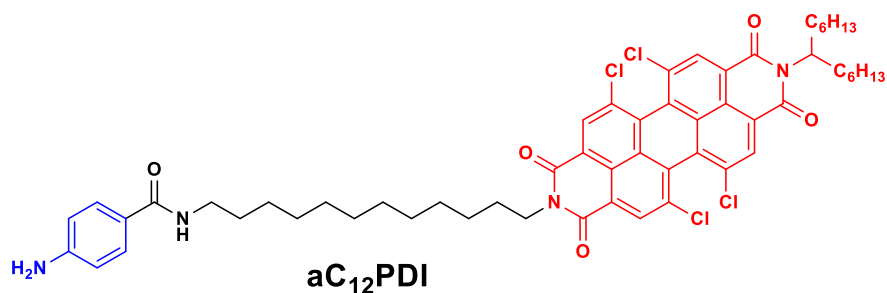


Figure 1. Structure of the amine used for in-situ grafting, the molecule denoted as *aC₁₂PDI*. The molecule consists of PDI (in red) and aniline (in blue) derivatives connected through an aliphatic 12-carbon chain. See Supplementary Information (SI-1) for details about synthesis.

2. Experimental section

Materials and reagents

The *aC₁₂PDI* compound was synthesized by a two-step procedure described in detail in SI-1. The final product *aC₁₂PDI* was purified in silica gel column chromatography and used for in-situ generation of diazonium salts, further called *dC₁₂PDI*. Tetrabutylammonium hexafluorophosphate (TBAPF₆ - Sigma-Aldrich), dichloromethane (DCM – HPLC grade, VWR chemicals), acetonitrile (ACN-HPLC grade, Carlo Erba), tert-Butyl nitrite (tBuONO – Sigma-Aldrich) were used as received. Milli-Q water (18 MΩ×cm) was used during polishing and for rinsing of the electrodes. A circular glassy carbon (GC) electrode was supplied by Bioanalytical Systems Inc. (Model MF-2012 with a surface area of 0.07 cm²).

Electrochemical grafting

Prior to each experiment, the surface of GC working electrode was polished with a 0.04 μm Al₂O₃ slurry (PRESI), sonicated in water and DCM for 2 min and rinsed with DCM. A BioLogic potentiostat (VSP model with 4 channels) was employed to carry out all electrochemical experiments. Electrochemical grafting was performed in a 3-electrode cell with

2 ml of 0.1 M TBAPF₆ DCM stagnant solution with 1 mM of aC₁₂PDI. Five equivalents of tBuONO were added into solutions and stirred for 10 min before imposing the grafting potential or the potential sweep for the in-situ generation of dC₁₂PDI from amine as described in literature.¹⁹ Note that the tBuONO compound is not electrochemically active (SI-2) and does not contribute to the electrochemical response. A Pt rod was used as a counter electrode and a Ag/AgNO₃ (0.01 M in 0.1 M TBAPF₆ ACN) was used as a reference electrode. An AC frequency of 10 kHz with 30 mV peak-to-peak amplitude was imposed on DC potential during chronoamperometric grafting experiments to follow evolution of capacitive impedance. Then, the capacitive impedance values were converted into surface coverage values from the calibration curves, providing kinetics of grafting as described in detail in ref.²⁰ and SI-3.

After the deposition of C₁₂PDI layer, the GC electrodes were sonicated for 2 min and washed with DCM. The surface coverage of PDI electroactive groups was estimated from cyclic voltammograms at 10 mV/s in a 0.1 M TBAPF₆ DCM solution according to the formula $\Gamma = Q/nFS$, where Γ is the surface coverage, Q is the charge estimated by integration of reduction peaks in CV, $n = 2$ is the number of electrons involved, $F = 96485$ C/mol is the Faraday constant and $S = 0.07$ cm² is the surface area of electrode.

All experiments with potential imposed more negative than -0.5 V were carried out in a glovebox (O₂ < 100 vpm and H₂O < 10 vpm) to avoid any interference of oxygen reduction process. All experiments were performed at room temperature of ca. 290 K.

Spectroelectrochemical characterization

Spectrophotometric measurements were carried out in a reflection mode using an in-house made bench composed of Princeton Instruments modules coupled to Biologic SP-150 potentiostat. Details of the instrumentation can be found elsewhere.^{21,22} In short, a 3-electrode cell with GC as working electrode, Pt as counter electrode and Ag/AgNO₃ (0.01 M in 0.1 M

TBAPF₆ ACN) as reference electrode (identical to electrochemical grafting) were employed. The connection between the light source, the electrochemical cell and the spectrophotometer was ensured through a “Y-shaped” optical fiber bundle: 18 fibers to guide the light to the cell, and 19 fibers to collect the reflected light from the cell directing it to the CDD detector (320-1080 nm, 0.5 nm spectral resolution, 2 MHz acquisition bandwidth). The acquisition rate of spectrometer was set to 10 frames per second that converted to 1 frame per mV at 10 mV/s.

EC-Lab software was used to control the BioLogic potentiostat and trigger the acquisition of spectral data, which guaranteed identical time scales of two independent measurements. In-house-built software was used for data plotting and processing. For each dataset, the reference intensity (I_{ref}) was obtained by averaging the intensity of the first 10 measured frames. Consequently, an absorbance variation (ΔAbs) and not an absorbance value (Abs) was monitored:

$$\Delta Abs = Abs - Abs_{ref} = -\log_{10} \left(I / I_{ref} \right) \quad (1)$$

Numerical simulations

The numerical simulations were carried out using KISSA-1D[©] software, version 1.2.2b.^{23,24} The Butler-Volmer relationship was used to describe all electrochemical reactions. The monolayer formation was described with the Langmuir isotherm. The formation of multilayers was described with a first-order adsorption kinetics achieved by setting the saturation surface coverage for the Langmuir isotherm in KISSA-1D[©] software 10^5 times larger in comparison to the value for the monolayer (*vide infra*). Simulations were performed for a planar electrode of 0.07 cm^2 under natural convection conditions with a stagnant layer thickness of $250 \text{ }\mu\text{m}$. Diffusion coefficients of all species were $5 \times 10^{-6} \text{ cm}^2 \text{ s}^{-1}$. The values of equilibrium potentials, kinetic constants, charge transfer coefficients and surface coverage at saturation are

summarized in Table 1. We also provide all technical details of KISSA-1D[®] simulations in SI-5.

3. Results and Discussion

Overview of dC₁₂PDI reduction

We performed the cycling voltammetry of aC₁₂PDI and in-situ generated dC₁₂PDI compounds in a large potential window to identify the onset potentials and reduction kinetics of PDI derivatives and diazonium salt.

Fig. 2a shows a typical voltammogram of aC₁₂PDI compound. On the cathodic sweep, the current is close to zero from 0.8 V to ca. -0.55 V then decreases showing two reduction peaks with later slight increase, reaching the diffusion-limited current of -50 $\mu\text{A cm}^{-2}$ at -1.3 V. Two oxidation peaks appear on the reverse scan with current gradually dropping to zero from ca. -0.55 V to 0.8 V. The half-way potentials for the first and the second reduction/oxidation steps are -0.66 V and -0.86 V correspondingly. In both cases, the peak-to-peak separation is ca. 80 mV, indicating the fast electron transfer typical for PDI derivatives.^{25,26}

The shape of voltammogram changes when the formation of dC₁₂PDI is initiated by addition of tBuONO in the solution of aC₁₂PDI (Fig. 2b). The two peaks, associated with the reversible reduction of PDI group, remain present in Fig. 2b. The position of these peaks and the peak-to-peak separation are identical to the ones observed on the voltammogram recorded on aC₁₂PDI compound in Fig. 2a. The absolute values of currents associated with PDI reduction and oxidation increase with cycling. Apart from the response of the PDI group, a slight decrease of the cathodic current at ca. 0.25 V is observed, reaching a plateau of ca. -2.5 $\mu\text{A cm}^{-2}$ between 0.05 V to -0.3 V (Fig. 2b inset) that is associated with the reduction of diazonium moiety.¹ This current drop is apparent only on the first cycle suggesting the suppression of diazonium reduction after the initial C₁₂PDI layer was formed.

Another new distinct feature is the appearance of the reduction peak at ca. -0.45 V with the current reaching of ca. $-30 \mu\text{A cm}^{-2}$. The position of this peak slightly shifts to more negative potentials with a slight increase of the reduction current on the consequent scans. Remarkably, on the anodic scans the current drops down to a negative value of ca. $-7 \mu\text{A cm}^{-2}$ at ca. -0.45 V, indicating a reduction process ongoing in this region. The origin of this peak is not clear. Note that the voltammogram recorded on C₁₂PDI layer (Fig. 2c) shows minor contribution of grafted PDI moieties to current density at -0.45 V in Fig. 2b. The position of reduction/oxidation peaks (Fig. 2c) lies more negative than -0.45 V and is similar to half-way potentials of aC₁₂PDI in Fig. 2a.

Below, we also provide experiments on the grafting at fixed potential as well as spectroelectrochemical studies to clarify the origin of the peak at -0.45 V in Fig. 2b and the mechanisms of C₁₂PDI layers formation.

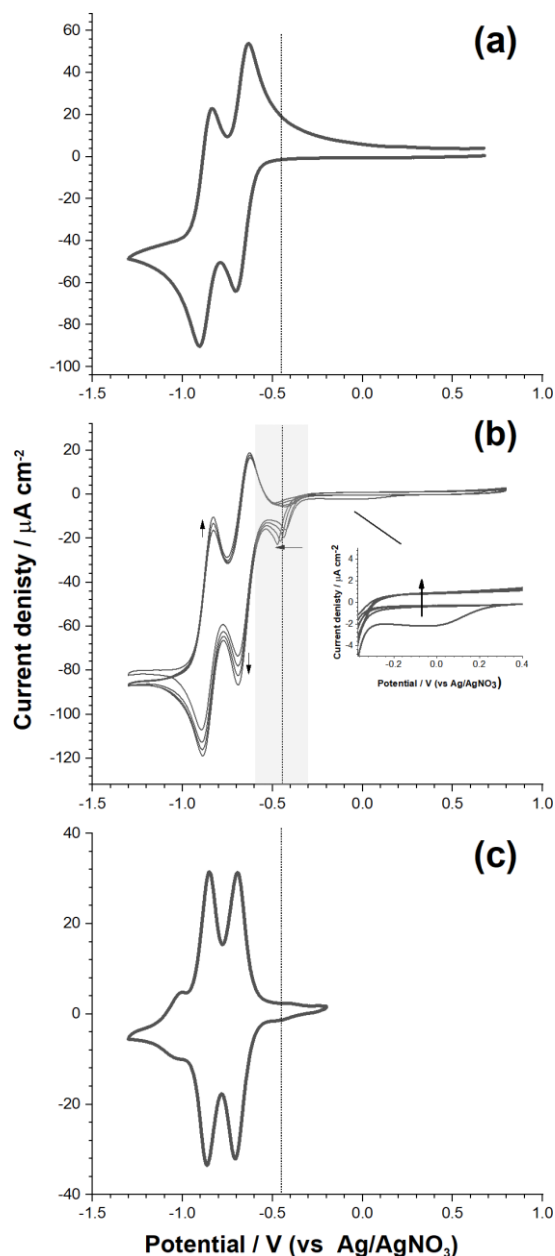


Figure 2. (a) Cyclic voltammograms of (a) 1 mM aC₁₂PDI, (b) 1mM aC₁₂PDI + 5 eq. of tBuONO and (c) grafted C₁₂PDI layer formed after 3 polarization cycles in figure (b) with a surface coverage of $24 \times 10^{-10} \text{ mol cm}^{-2}$. Experiments were performed on GC in a 0.1 M TBAPF₆ DCM solution in a glovebox. The potential sweeps start at the most positive potential, progresses towards -1.3 V and returns to the initial potential at 10 mV/s. The dotted lines are placed at the fixed potential of -0.45 V and semitransparent grey box covers the potential range of -0.3 to -0.6 V for the visual reference.

Surface coverage of C₁₂PDI layer as a function of applied potential

Evolution of the surface coverage of C₁₂PDI layers formed on GC under applied potential in Fig. 3a shows 4 distinct regions.

In the 1st region, ranging from 0.3 V to 0.8 V, the surface coverage is close to zero and no grafting occurs.

At ca. 0.25 V, the grafting starts with $\Gamma = 1.2 \times 10^{-10}$ mol cm⁻² after 15 min of grafting and $\Gamma = 2.4 \times 10^{-10}$ mol cm⁻² after 30 min of grafting. As the potential decreases down to -0.3 V, the surface coverage gradually increases up to $\Gamma = 3.0 \times 10^{-10}$ mol cm⁻² for both, 15 min and 30 min of grafting. This value is close to the surface coverages observed for PDI-based SAMs²⁵ suggesting the formation of C₁₂PDI monolayers. Kinetics of layer grafting in the region of -0.3 V – +0.25 V shows a Langmuir type behavior (Fig. 3b, curve at 0 V) with a rapid increase of surface coverage during initial 100 - 200 s, followed by the gradual transition to a steady-state value of ca. 3.0×10^{-10} mol cm⁻². Importantly, the surface coverage did not exceed this value, even after 2h grafting at 0 V which is not the case for grafting of other diazonium salts of benzene derivatives (nitrobenzene, carboxybenzene etc)²⁷⁻²⁹. Conventionally, a grafting of multilayers for mentioned molecules is explained by slow electron transfer through grafted layer, reduction of aryl diazonium to aryl radical and its grafting on the preformed layer.²⁷ The key difference between mentioned benzene derivatives and dC₁₂PDI compound is the presence of the aliphatic 12-carbon chain, which supposedly blocks completely the electron transfer *via* C₁₂PDI monolayer and prevents the diazonium reduction and multilayer formation.³⁰

The 3rd region covers the potential range between -0.35 V and -0.55 V where the surface coverage significantly overpasses the characteristic value of a monolayer, reaching the maximum at ca. -0.4 V. Remarkably, this region coincides with the extension of peak at ca. -0.45 V in Fig. 2b, marked with a semitransparent grey box. The absolute values of surface

coverage depend on the grafting time. For example, at -0.4 V, Γ equals 36×10^{-10} mol cm⁻² after 15 min of grafting and it increases to 62×10^{-10} mol cm⁻² after 30 min of grafting. We also performed a 2 h long experiment at -0.4 V and were able to reach $\Gamma = 10^{-8}$ mol cm⁻². In contrast to the Langmuir type kinetics at 0 V, the surface coverage increases linearly with time in this region, as illustrated on the example of grafting at -0.5 V in Fig. 3b, demonstrating a first order kinetics.

All these experimental observations suggest that, despite the presence of a 12-carbon chain separator between the PDI moiety and the phenyl group, the C₁₂PDI layer supports the diazonium reduction in the potential range of -0.35 V and -0.55 V in contrast to -0.3 V – +0.25 V. The previous studies on grafted nitrobenzene-, benzoylbenzene- and polyphenylene-layers^{18,31} reported that the reduction of deposited layers promotes electron transfer *via* the tunnel effect through organic layer,³⁰ that can be the case of the C₁₂PDI layer too. The survey of the redox state of C₁₂PDI layer will be performed in the next section.

At potentials more negative than -0.6 V, the surface coverage abruptly drops to values close to zero (Fig. 3a), coherent with grafting kinetics in Fig. 3b. Similar effect of surface coverage decrease at cathodic overpotentials was observed by adding redox mediators in the grafting solution (DPPH¹⁴ and O₂¹⁶). In our case, the only redox mediator present in the solution is the PDI entity of aC₁₂PDI and dC₁₂PDI molecules. Seemingly, the subtle differences in the reduction kinetics of PDI moieties present in the solution and grafted on GC interface play a crucial role in the mechanisms of diazonium grafting. In the next section, we provide comparative spectroelectrochemical study to unravel these effects.

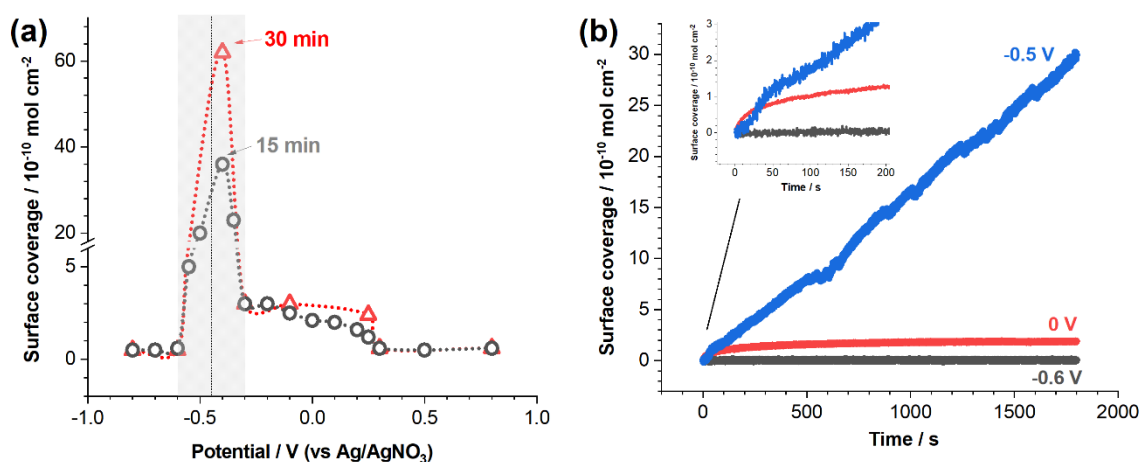


Figure 3. (a) Surface coverage of $C_{12}PDI$ layers as a function of applied potential during electrodeposition. The potential was imposed for 15 min (grey circles) and 30 min (red triangles) using GC working electrode and 1 mM $aC_{12}PDI + 5eq. of tBuONO$ in 0.1 M $TBAPF_6$ DCM. The dotted line is placed at the fixed potential of $-0.45 V$ and semitransparent grey box covers the potential range of -0.3 to $-0.6 V$ for the visual reference. (b) Surface coverage vs time defined from in-situ capacitive measurements²⁰ at fixed potentials of 0 V, $-0.5 V$ and $-0.6 V$. All experiments at potentials more negative than $-0.5 V$ were performed in a glovebox.

Reduction states of PDI moieties in solution and in $C_{12}PDI$ layer as a function of potential

Fig. 4 shows the evolution of relative absorbance of PDI moiety in the grafted $C_{12}PDI$ layer on GC during a polarization cycle. The voltage ramp (Y axis) starts at $-0.2 V$ then it progresses towards $-1.3 V$ (horizontal blue line in Fig. 4) and returns to the initial potential of $-0.2 V$. The axial symmetry along a $-1.3 V$ line implies that the changes of the electronic structure of PDI moiety are reversible and do not depend on the direction of polarization. For this reason, we consider only the bottom part of spectral signature that shows the first part of the polarization cycle from $-0.2 V$ to $-1.3 V$. Note that ΔAbs values ($\Delta Abs = Abs - Abs_{ref}$, with Abs_{ref} = absorbance recorded during the first 10 frames, see Eq. 1 in experimental section) are measured. Consequently, an electrochemical consumption of PDI moiety (band at 523 nm) results in

negative absorbance values and the generation of PDI^- (at 785 nm) and PDI^{2-} (at 693 nm) from redox reaction leads to positive values of absorbance.

The subtle change of color in Fig. 4 of PDI^- band from -0.2 V to -0.6 V indicates the onset of PDI reduction at ca. -0.4 (Fig. 4). This coincides with the decrease of current density, which, therefore, can be attributed to a faradaic process of PDI^- formation. As potential decreases in Fig. 4, the values of PDI^- and PDI^{2-} absorbance start to increase coherent with the appearance of a second peak on current density curve. Remarkably, there is no variation in absorbance at ca. -1.0 V where there is a small peak on the current density curve. This suggests that the source of current related to this peak is not related to faradaic processes of PDI reduction but probably is due to ion migration within the PDI layer¹⁸ and the layer reorganization due to the electrostatic repulsion between negatively charged PDI^{2-} groups etc.

Comparison of PDI moieties reduction (in solution and within the grafted layer) in Fig. 5 shows that the PDI in the bulk solution undergoes reduction at ca. -0.55 V, which is ca. 0.2 V more negative than PDI in the grafted layer. The potential window of selective PDI reduction in C_{12}PDI layer is coherent with the potential window of multilayer formation in Fig. 3a and the extent of the peak at ca. -0.4 V in voltammograms of dC_{12}PDI compound in Fig. 2b (highlighted with semitransparent grey box).

Overall, the correlation of spectroscopical and electrochemical results around -0.45 V suggests that the reduction of deposited PDI moieties promotes tunnel effect through an organic layer and therefore, facilitates the electron transfer *via* the formed layer. In this vein, the peak in current density at -0.45 V in Fig. 2b corresponds to an enhanced reduction of diazoniums to aryl radicals. Once generated, highly reactive aryl radicals can chemically bond to C_{12}PDI layer with the formation of multilayers. Below, we summarize the proposed reaction sequence in the whole potential range and provide numerical simulations to reinforce our findings.

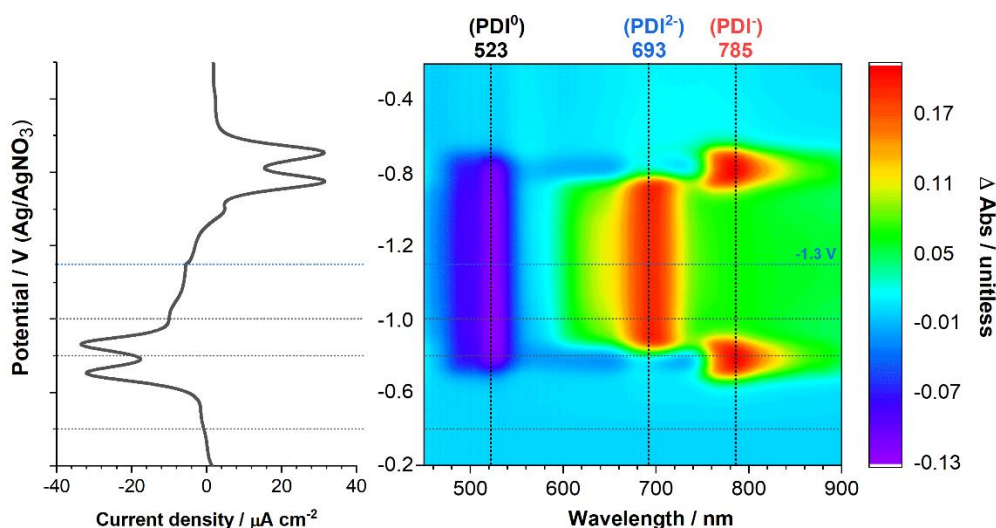


Figure 4. Absorbance variation (on the right) of grafted $C_{12}PDI$ layer on GC ($\Gamma = 24 \times 10^{-10}$ mol cm^{-2}) in 0.1 M $TBAPF_6$ DCM under a polarization cycle between -0.2 V to -1.3 V at 10 mV/s. The variation of simultaneously recorded current is presented on the left. The experiments were performed in a glovebox. The dotted lines indicate the bands positions attributed to the conversion of PDI^0 to PDI^- and PDI^{2-} .^{25,26,32}

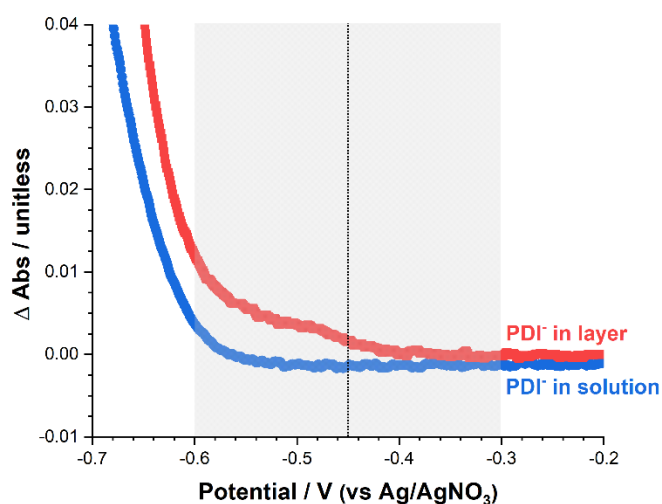
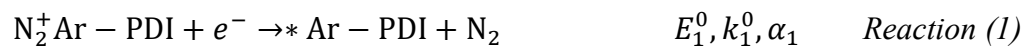


Figure 5. Comparison of absorption change values for PDI^- species in the grafted layer and under thin-layer conditions around the onset of PDI reduction. Details on experiment in the bulk solution are provided in SI-4. The dotted line is placed at the fixed potential of -0.45 V and semitransparent grey box covers the potential range of -0.3 to -0.6 V for the visual reference.

Tentative mechanism of dC₁₂PDI grafting

The exact mechanism of diazonium grafting that would encompass all aspects of deposition is still a subject of ongoing research, even in the case of a relatively simple and more studied NBD molecule.^{33,34} It is therefore not reasonable to propose a holistic mechanism of dC₁₂PDI grafting but we rather focus on building the phenomenological model that would qualitatively describe the observed relationship between surface coverage, grafting kinetics and applied potential. Numerical simulations were carried out in Kissa-1D[©] software (see Experimental section and SI-5 for details).

The electrografting process can be described by one-electron transfer reaction and two irreversible first order reactions:^{17,33,34} the reduction of diazonium (dC₁₂PDI noted as N₂⁺Ar – PDI) with the formation of aryl radical (* Ar – PDI) (reaction 1)^{35,36}, followed by grafting of aryl radical to the GC surface with the formation of a covalent bond (reaction 2)^{37,38} or consumption of highly reactive aryl radicals *via* side reactions involving solvent or other species (reaction 3).^{35,39} The diazonium reduction was considered irreversible in accordance with experimental data in this work.



The grafted layer undergoes reversible reduction at potentials lower than ca. -0.4 V. The deposition of non-reduced layer (reaction 2) follows Langmuir adsorption isotherm but as soon as the deposited layer is reduced, the formation of multilayers follows *via* enhancement of tunnel effect as discussed above. The multilayer growth is known to be a complex process that should at least depend on heterogeneous electron transfer rate that in turn depends on

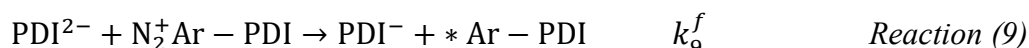
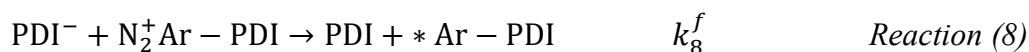
dynamically changing structure of deposited layer.^{33,34} For the sake of simplicity, we neglect these complex effects and consider the layer reduction ongoing *via* two reversible electron charge transfer in reactions 4 and 5. The formation of a multilayer was described with a first order kinetics, coherent with our experimental data (*vide supra*).



PDI moieties present in solution also undergo reduction with the onset potential of ca. 0.2 V more negative in comparison to grafted PDI groups. The source of PDI moieties can equally originate from dC₁₂PDI, aC₁₂PDI or side products produced *via* reaction 3. We do not make any distinction between these species, including all of them *via* reactions 6 and 7.



The key moment of the mechanistic scheme is the inclusion of irreversible cross-reactions between reduced PDI moieties and diazonium compounds in the solution *via* reactions 8 and 9. Therefore, diazonium can be reduced in the bulk solution and resulting aryl radicals can be scavenged through formation in side products *via* reaction 3. This restricts the concentration of aryl radicals in vicinity of GC interface and inhibits the layer deposition.



The input parameters for reactions 1-9 were either taken from literature or estimated from our experimental data and are summarized in Table 1. Results of numerical simulations are presented in Fig. 6.

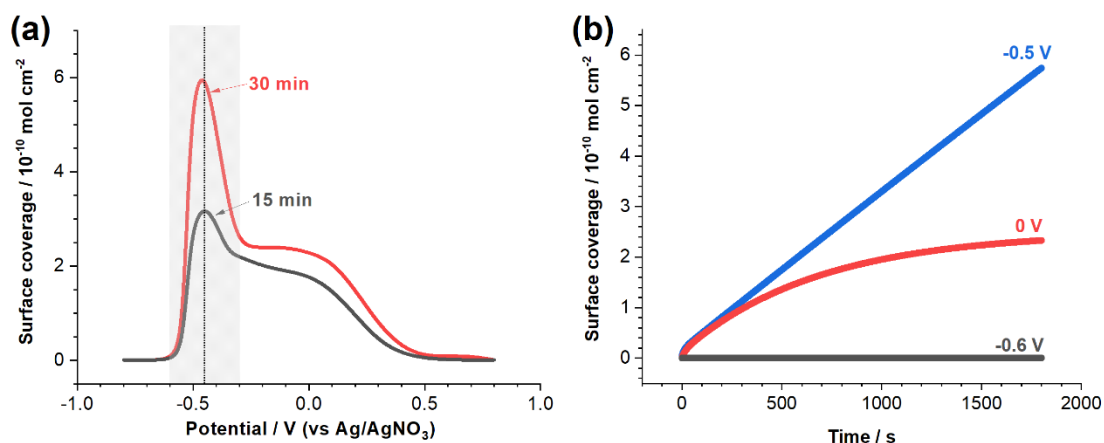


Figure 6. Results of numerical simulations in Kissa 1D[®] software of reactions 1-9 with input parameters in Table 1. (a) shows the surface coverage of deposited C₁₂PDI layers as a function of applied potential. The potential was imposed for 15 min (grey curve) and 30 min (red curve). The dotted line is placed at the fixed potential of -0.45 V and semitransparent grey box covers the potential range of -0.3 to -0.6 V for the visual reference. (b) Surface coverage vs time at fixed potentials of 0 V, -0.5 V and -0.6 V.

There is a good agreement between simulated and experimental data of surface coverage vs applied potential and grafting kinetics (compare Fig. 3 and Fig. 6). This validates the proposed mechanistic sequence and in particular the role of PDI moiety in first, favoring the formation of multilayer in a small potential window of ca. -0.4 V to ca. -0.6 V and second, inhibition of a layer deposition through cross-redox reactions at potentials lower than -0.6 V in the bulk solution.

The biggest difference between experimental and simulated data concerns the absolute values of surface coverage related to the multilayer formation (marked with the semitransparent grey box). Obviously, this is related to the simplifications imposed on the model in terms of multilayer formation. Likewise, the structure of multilayer formation and its impact on heterogeneous electron transfer were not considered. Quantitative description of these effects

requires detailed in-situ spectroelectrochemical characterization of C₁₂PDI multilayers and will be the subject of our future work.

Table 1. Input parameters of numerical simulations where E^0 is the standard potential, α is the transfer coefficient, Γ^s is the s surface coverage at saturation and k is the rate constant (k^0 – standard electrochemical rate constant, k^{ad} – adsorption rate constant, k^f – forward rate constant). The reference is given for any value taken from the literature; otherwise, the values were estimated from experimental data in this work. The initial concentration of all species was zero except for 1 mM N₂⁺Ar – PDI and 1 mM PDI.

Reaction	E^0	$k(k^0, k^{ad} \text{ or } k^f)$	α	Γ^s
1	-0.10 V	0.003 cm s ⁻¹ ³⁶	0.3 ³⁶	-
2	-	2×10 ⁵ M ⁻¹ s ⁻¹	-	2.5×10 ⁻¹⁰ mol cm ⁻²
3	-	10 ⁸ s ⁻¹ ^{35,39}	-	-
4	-0.40 V	100 s ⁻¹	0.5	10 ⁻⁵ mol cm ^{-2*}
5	-0.60 V	100 s ⁻¹	0.5	10 ⁻⁵ mol cm ^{-2*}
6	-0.66 V	100 cm s ⁻¹	0.5	-
7	-0.86 V	100 cm s ⁻¹	0.5	-
8	-	10 ⁸ s ⁻¹ ¹⁷	-	-
9	-	10 ⁸ s ⁻¹ ¹⁷	-	-

*The values of Γ^s for multilayer formation were set to an arbitrary large value to provide a first order deposition kinetics in the Langmuir formalism uniquely available in KISSA-1D[©] interface.

4. Conclusion

This study has elucidated the grafting mechanism of dC₁₂PDI molecule as a function of applied potential. Three regimes, based on the difference in grafting kinetics have been

highlighted for C₁₂PDI layer formation. The first regime comprises grafting, which obeys Langmuir type kinetics and is restricted to the deposition of monolayers in the potential range above PDI moieties reduction. The second regime consists of grafting multilayers following the first order kinetics in the potential range restricted to the selective reduction of PDI moieties in the deposited layer (but not in the bulk solution). When species containing PDI groups undergo reduction in the solution, the deposition of C₁₂PDI layer is strongly inhibited. The observed behavior was explained by the redox cross-reaction involving PDI reduced moieties in the bulk solution, and further reinforced with numerical modeling. This study demonstrates that the fine control of grafting kinetics of the redox active PDI moiety can be achieved through tuning the applied potential that, in principle, can be extended to any reducible active group.

Supporting Information

SI-1 Synthesis and characterization of aC₁₂PDI compound; SI-2 Electrochemical response of tBuONO solution; SI-3 Methodology of in-situ estimation of surface coverages via capacitive impedance measurements; SI-4 Spectroelectrochemical signature of aC₁₂PDI in the bulk solution; SI-5 Input parameters in KISSA-1D© interface (v. 1.2.2b).

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TOC Graphic

