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Title

In situ formation of rare-earth-doped nanoparticles in a silica matrix from Molecular Dynamics simulations

Authors

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Abstract

Developing new rare-earth-doped optical glasses with “enhanced” spectroscopic properties requires the elaboration of new glass compositions. To overcome some typical limitations of silica glass, a strategy consists in encapsulating rare-earth (RE) ions within oxide nanoparticles (NPs) through a phase separation mechanism. In this paper, Molecular Dynamics simulations were performed using an interatomic potential reproducing the phase separation within a MgO-SiO₂ binary melt to obtain RE-codoped glass models with RE = Eu or Er. In these structures, we observed that Mg-rich regions, included into a silica-rich matrix and identified as NPs, are amorphous and exhibit a large range of sizes. We showed that such nanoparticles are the host of a depolymerization phenomenon of the NPs' SiO₄ tetrahedral network leading to a release of non-bridging oxygen atoms. In the NPs, the MgO concentration increases and the doping RE ions are mainly located into the NPs where they are over-concentrated compared with the nominal doping concentration. However, the induced clustering effect is limited because of the non-bridging-oxygen-rich environment encountered in the NPs. This numerical analysis allows to give an insight on the chemical composition of the NPs, and especially on the local environment of the encapsulated rare-earth ions.

Keywords

Molecular Dynamics ; Nanoparticles ; MgO ; Silicate ; Rare-earth ; Glass.

Main text

1. Introduction

Rare-earth (RE)-doped silica-based glasses are materials of great importance for technological applications such as fiber lasers, optical amplifiers, sensors..., and have grown considerably in recent years. They combine luminescent properties of RE ions with the chemical, mechanical and financial advantages of the silica glass. However, the development of new applications using them as an amplifying material through the stimulated emission of the luminescent ions requires rethinking the nature and structure of these doped glasses to achieve “augmented” spectroscopic properties [1, 2], which would not appear in a pure silica local environment. To develop such new silica-based glasses, a strategy consists in incorporating RE-doped oxide nanoparticles (NPs) in a silica glassy matrix [3]. The encapsulation of luminescent ions in NPs enables to practice engineering of the luminescent response of the material through the control of the structure and composition of the RE environment [4, 5]. To obtain RE-doped NPs in silica matrix, Blanc *et al.* [6] take advantage of the so-called phase separation phenomenon [7] occurring at high temperature in the MgO-SiO₂ binary system during melt/quench sequences of glass fabrication. After quenching process, phase separation leads to a binary glass composed of two different phases: a Mg-rich Si-poor phase, identified as the NP, and a Si-rich Mg-poor phase, identified as the silica host matrix. Secondary Ion Mass Spectroscopy (SIMS) and Atom Probe Tomography (APT) measurements have confirmed the Mg-rich composition of such NPs wherein luminescent ions are mainly located [8, 9]. Vermillac *et al.* [10] have highlighted a broadening of the emission band of the ⁴I_{13/2} level of Er³⁺ in erbium-doped silica-based fibers that were co-doped with magnesium or lanthanum. Recently, Blanc *et al.* [9] have shown, for the first time, a NP size-dependence of the chemical composition of amorphous Er-doped Mg-silicate NPs produced *in situ* in a silica glass. However, detailed structural properties such as RE local environment, coordination numbers, oxygen species, are difficult to obtain from experiments because of the size of the NPs and of the low concentration of RE used for doping. Therefore, numerical simulations are very helpful to further investigate the structural properties at the nanoscale.

The aim of this work is to discuss in detail the structural and compositional properties of RE-doped Mg-silicate NPs as well as the local environment encountered by RE ions. To this purpose, Molecular Dynamics (MD) simulations have been performed using an interatomic adaptive potential developed by Bidault *et al.* [11] that has allowed to reproduce the phase separation of the $x\text{MgO}-(1-x)\text{SiO}_2$ binary glass. Through melt/quench protocol, we have elaborated three glasses with different chemical compositions: $0.10\text{MgO}-0.90\text{SiO}_2$, $0.01\text{ErO}_{3/2}-0.10\text{MgO}-0.89\text{SiO}_2$, and $0.01\text{EuO}_{3/2}-0.10\text{MgO}-0.89\text{SiO}_2$. From these modeled structures, the RE-doped Mg-silicate NPs are identified and characterized in size, shape, and composition. The structure of magnesium and RE environments is analyzed in details, and we particularly focus on the effect of the Mg codoping on the RE local environment. The question of the RE clustering is finally discussed.

2. Material and methods

2.1. Interatomic potential

Since silica is a glass former of prime importance, a large variety of interatomic potentials has been designed to investigate its structural and dynamical properties [12-16]. Among them, pairwise potentials with partial charges have been extensively used, since they allow, for a reasonable calculation cost, to model glass samples containing up to several hundreds of thousands atoms during millions of timesteps. In this work, we adopt a modified version of the interatomic pairwise potential developed by Pedone *et al.* [12], which is known to accurately reproduce the structural properties of amorphous SiO_2 . Its functional form is a sum of a long-range Coulomb term, a short-range Morse function, and a repulsive contribution. Thus, the potential between particles i and j is given by:

$$U(r_{ij}) = \frac{q_i q_j}{r_{ij}} + D_{ij} \left[\left\{ 1 - e^{-a_{ij}(r_{ij}-r_0)} \right\}^2 - 1 \right] + \frac{C_{ij}}{r_{ij}^{12}} \quad (1)$$

where q_i and q_j are partial charges, D_{ij} , C_{ij} , a_{ij} and r_0 are adjustable parameters and r_{ij} is the interatomic distance. Short-range interactions were truncated at 5.5 \AA . The Coulomb part was evaluated using Wolf method [17] with a long-range cutoff set as 7.5 \AA and a damping parameter set as 0.30 \AA^{-1} .

In a previous work, Bidault *et al.* [11] highlighted the inability of a fixed charge model to reproduce the phase separation phenomenon occurring in binary silicate glasses. The authors refined the Pedone potential in order to reproduce the phase diagram of the $\text{MgO}-\text{SiO}_2$ system, by taking into account the different ionicity of Mg-O and Si-O bonds, which is known to be crucial to model the phase separation

phenomenon [18]. In order to respect the electroneutrality of the binary system and to explore the whole phase diagram by means of a unique set of parameters, each oxygen charge of the O-O repulsive interaction dynamically reacts to its own cationic environment. Details on the procedure can be found in Ref. [11]. These modifications allow to reproduce the separation of large phases in the MgO-SiO₂ binary system: Mg-rich Si-poor ones on one hand, and Si-rich Mg-poor ones on the other hand. Since phase separation phenomenon occurs in REO_{3/2}-SiO₂ systems (with RE = Eu and Er) [18, 19], it is also necessary to consider the ionicity of Er-O and Eu-O bonds. All potential parameters used in this study are listed in Table 1.

[Tab 1 around here]

2.2. Simulation procedure

MD simulations were performed to obtain three glass samples with a volume of approximately 200×50×200 Å³ and molar compositions of 0.10MgO-0.90SiO₂ (150336 atoms: 5184 Mg), 0.01EuO_{3/2}-0.10MgO-0.89SiO₂, and 0.01ErO_{3/2}-0.10MgO-0.89SiO₂ (151824 atoms: 5248 Mg, and 544 Eu or Er). They were carried out by the LAMMPS (Large-scale Atomic/Molecular Massively Parallel Simulator) program [20] with periodic boundary conditions. The Velocity-Störmer-Verlet algorithm was used to solve Newton's equations of motion with a 1 fs timestep.

Glass sample elaboration was done at constant number of atoms, pressure and temperature using a Berendsen barostat and a Langevin thermostat. For each of the three glasses, the same numerical protocol was applied: in a first step, a silica glass was obtained by melting and equilibrating a cristobalite silica structure at 4000 K during 0.1 ns. Then, the sample was cooled down to 300 K at a quench rate of 5×10¹² K.s⁻¹. Secondly, O, Mg, and RE (Eu or Er) atoms were placed randomly in the silica matrix to obtain the desired chemical composition. Then, another melt/quench sequence was performed in order to obtain the final glass model: the system was melted and equilibrated at 4000 K during 0.1 ns, and quenched at 5×10¹¹ K.s⁻¹ from 4000 K to 2400 K; then, long stages of 6.4 ns each were successively performed at 2400 K, 2300 K, 2200 K, 2100 K, 2000 K, and 1900 K, a temperature range wherein the phase separation occurs [7]; at 1900 K, which is below the glass transition temperature T_g, the system is almost fixed (for this quench rate, T_g = 2210 K, 2150 K, 2160 K for 0.10MgO-0.90SiO₂ glass, 0.01ErO_{3/2}-0.10MgO-0.89SiO₂ glass, and 0.01EuO_{3/2}-0.10MgO-0.89SiO₂ glass, respectively). Finally, the system was quenched at 5×10¹¹ K.s⁻¹ from 1900 K to 300 K. All the results presented in this paper were derived from these three equilibrated structures at 300 K and atmospheric pressure.

3. Results and Discussion

3.1. Nanoparticles' characteristics

MD visualizations were obtained using OVITO software [21]. In Fig. 1a, snapshots of the 0.10MgO-0.90SiO₂ glass structure obtained by MD show the existence of two different kinds of region: Mg-rich Si-poor phases on one hand and Si-rich Mg-poor phases on the other hand. This observation of compositional heterogeneities is in agreement with experimental analysis using Secondary Ion Mass Spectroscopy (SIMS) and Atom Probe Tomography (APT) techniques [8, 9], as observed in Fig. 1b, where NPs seem to be amorphous and have a non-spherical shape. Figure 2 displays the Mg-Mg Radial Distribution Function (RDF) and Mg-Mg Cumulative Distribution Function (CDF) for the three glass models. We observe that the Mg-Mg RDF first peak is largely broadened around a mean distance of 3.0 Å, indicating that there is no crystalline arrangement in the MgO rich phases.

[Fig 1 & 2 around here]

Fig. 3 displays snapshots of the undoped and the RE-doped glass structures. Only Mg atoms are represented with a color code related to the size of the Mg-rich phase in which they are located. Two Mg atoms separated by less than 4.4 Å (this Mg-Mg cut-off distance corresponds to the position of the first minimum of the Mg-Mg RDF, see Fig. 2) are considered to belong to the same Mg-rich phase. We now identified these Mg-rich regions as nanoparticles (NPs). A visual analysis of snapshots indicates that NPs have a non-spherical shape and are inhomogeneously distributed in the silica matrix. They also exhibit a wide range of sizes: for the three glass models, about 10 % of Mg atoms are isolated in silica, and about 60 % are involved into NPs composed of 28 Mg atoms and more. The largest NP observed among these three glass samples is composed of 525 Mg atoms, with a length of about 100 Å.

[Fig 3 around here]

[Tab 2 around here]

Table 2 reports the NPs composition according to their size. NP composition is determined as following: an atom R (R = Si, Eu or Er) is considered belonging to a NP if it is bounded to an oxygen atom which is itself bounded to a Mg atom of the NP. In practice, an atom R belongs to a NP if it is found at a distance shorter than the first minimum of R-Mg RDF from a Mg atom of the NP. For the three glass

samples, we notice that the MgO concentration of a NP increases with its size. Whatever the NP size, rare-earth ions are over-concentrated compared to their nominal concentration of 1 %. Only 12.5 % of Eu ions (7.5 % of Er ions) are isolated in the silica matrix. While the $\text{EuO}_{3/2}$ concentration in the NPs seems to be almost constant whatever their size, the $\text{ErO}_{3/2}$ concentration significantly increases in the largest NPs.

3.2. Structure of magnesium environment

Mg^{2+} cation is known to act as a network modifier in silicate melts, according to several models describing most of the relevant aspects which lead to glass formation [22, 23]. To appreciate the tetrahedral connectivity, it is commonly used to distinguish five Q^n silicon species according to the number ($n = 0 \dots 4$) of bridging oxygen atoms per silicon tetrahedron. In such a way, a pure and perfect vitreous silica should exhibit 100% of Q^4 , meaning that all Si atoms are connected to four bridging oxygen (BO) [24]. Since the addition of MgO forms non-bridging oxygen (NBO), the Q^n distribution should become dependent on the MgO concentration. Table 3 reports Q^n distributions for Si atoms within the NPs of the 0.10MgO-0.90SiO₂ glass, as a function of the NPs size. We can observe that as the NP size increases, the fraction of Q^4 decreases, while the number of Q^2 and Q^3 units increases. As expected, the degree of depolymerization of the SiO₂ network in the NP increases with their size, which is due to the fact that the MgO concentration also increases with the NPs size.

[Tab 3 around here]

We now report and discuss the local environment of Mg atoms and its effect on the depolymerization of the silica network in the 0.10MgO-0.90SiO₂ glass model. Figure 4 displays the first peak of the Mg-O RDF as a function of the NP size in which Mg atoms are involved. We can observe a change in the shape and position of this peak. The average Mg-O nearest neighbor distance and coordination number increase from 1.92 Å ($N_{\text{Mg-O}} = 5.6$) for Mg in the silica matrix, to 2.04 Å ($N_{\text{Mg-O}} = 6.2$) in the largest NPs. It is also interesting to note that for an isolated Mg atom in the silica matrix, the Mg-O RDF appears to be the result of two main contributions, indicating that two different oxygen species are encountered: oxygen atoms bonded to only one Si atom, which may be the closest to Mg (corresponding to the first peak), and oxygen atoms bonded to two Si atoms, which may be the farther from Mg (corresponding to the shoulder after the first peak). This assumption is analyzed and verified hereafter.

[Fig 4 around here]

To investigate in more detail the Mg local environment, we use the terminology used by Afify and Mountjoy [25] to define the different oxygen species, as depicted on the right side of Fig. 5: a BO also bounded to one Mg atom is called BO*; a NBO bounded to one Mg atom is denoted NBO*, whereas a NBO bounded to two Mg atoms is denoted NBO**; finally, a Non-Network Oxygen (NNO) is an oxygen atom not bounded to any Si, but only to Mg atoms. The left side of Fig. 5 displays the Mg-O RDF and its contributions due to the previously defined oxygen species in the 0.10MgO-0.90SiO₂ glass model. We observe that the first maximum position is different for each contribution of the overall MgO RDF. It is located at around 1.90 Å, 2.05 Å, and 2.25 Å, for Mg-NBO*, Mg-NBO**, and Mg-BO*, respectively. The Mg-BO* interatomic distance is longer than the Mg-NBO ones, because a BO* is strongly bounded to two Si atoms. In the same way, Mg-NBO** distance is longer than the Mg-NBO* one, because a NBO** is constrained to be bounded to another Mg.

[Fig 5 around here]

Table 4 reports, for the same glass model, detailed oxygen coordination number of Mg as a function of the size of the NP in which it is located. The oxygen coordination number of Mg is provided by the area measured under the first peak of the RDF, using a cut-off distance of 2.9 Å (which is the position of the RDF first minimum). In this glass model, the average oxygen coordination number of Mg is found to be 6.0. But results reported in Table 4 show that this coordination number depends on the NP size in which Mg atoms are considered, increasing from 5.6 when isolated in the silica matrix to 6.2 in the largest NPs. These modeled coordination numbers are in agreement with experimental values obtained by ²⁵Mg 3QMAS NMR [26] for an enstatite glass and by ²⁵Mg NMR spectroscopy [27] for xMgO-(1-x)SiO₂ glasses where 0.50 < x < 0.67. These experimental results support the idea of a distorted octahedral environment for Mg. In our modeled structures, it is found that with the increase of the NPs size, the coordination of Mg with BO* and NBO* decreases, and consequently the oxygen-rich environment found in the largest NPs is mostly due to the presence of NBO** and NNO. Table 4 also reports the distribution of sites occupied by Mg atoms according to the number of coordinated oxygen atoms and as a function of NPs size. Isolated in the silica matrix, Mg atoms are mainly surrounded by 5 and 6 oxygen atoms. When the NP size increases, these 5- and 6-coordinated sites are less and less numerous, in favor of a majority of 6- and 7-coordinated sites for Mg in the largest NPs. Not reported here, the same analysis has been performed for the two RE-codoped glasses, and the same results on the Mg environment have been obtained. According to these results, the concentration in MgO and the average oxygen coordination of Mg are found to be nanoparticle size-

dependent. The Mg-O coordination number is mostly due to the presence of non-bridging-oxygen, which is in good agreement with the network modifier role of Mg^{2+} cation.

[Tab 4 around here]

3.3. Structure of rare-earth environment

From the two glass models of $0.01\text{ErO}_{3/2}\text{-}0.10\text{MgO}\text{-}0.89\text{SiO}_2$ and $0.01\text{EuO}_{3/2}\text{-}0.10\text{MgO}\text{-}0.89\text{SiO}_2$ composition, the structure of the rare-earth ion's environment has been analyzed in detail. Fig. 6 displays the RDF and CDF for the RE-Mg and RE-Si pairs (RE = Eu and Er). Eu-Mg and Er-Mg RDFs presents an intense first peak but also a well-defined second peak meaning that Mg atoms are also present in the second cationic coordination sphere of the rare-earth ions. Considering the CDF curves at a cut-off radius corresponding to the RDF first minimum (4.7 Å for Eu-Mg, 4.3 Å for Eu-Si, 4.6 Å for Er-Mg and 4.1 Å for Er-Si), one can deduce that the first cationic coordination sphere of Eu is on average composed of 2.3 Mg and 7.4 Si, while Er is surrounded on average by 3.2 Mg and 6.8 Si. Since it has been previously seen that NPs are Mg-rich environments, it can be concluded that RE ions are mainly located in the NPs. Moreover, considering the differences between Eu and Er, one can note that the environment of Er is richer in Mg than the one of Eu. Therefore, since the highest MgO concentration is found in largest NPs, it is expected to find a large fraction of Er^{3+} ions located in largest NPs. Actually, more than 40% of Er ions, against less than 25% of Eu ions, are found to be located in the largest NPs (composed of 126 Mg and more). This is in agreement with the previous analysis on the NP composition, where it has been shown that the largest NPs are composed of 3.3% of $\text{ErO}_{3/2}$ against 2.3% of $\text{EuO}_{3/2}$ according to the rare-earth doping. The origin of this difference is discussed hereafter in the paper.

[Fig 6 around here]

In order to study the oxygen coordination of the rare-earth ions and especially to understand the effect of the Mg codoping, two glass models of composition $0.01\text{EuO}_{3/2}\text{-}0.99\text{SiO}_2$ and $0.01\text{ErO}_{3/2}\text{-}0.99\text{SiO}_2$ have been produced for comparison. Fig. 7 displays the RDF and CDF for the Eu-O and Er-O pairs. It can be observed that Eu-O and Er-O interatomic distances are shifted to larger distances with the addition of Mg, which is the consequence of a change in the nature of the RE ions close environment. Using cut-off distances corresponding to the RDFs first minima (3.0 Å for Eu-O and 2.9 Å for Er-O), the CDF curves indicate oxygen coordination numbers increasing from 6.07 to 6.61 for Eu, and from 5.92 to 6.49 for Er. It means that for both RE ions, Mg codoping involves a significant increase of their oxygen coordination numbers.

[Fig 7 around here]

In order to understand and compare the structuration of the RE first coordination sphere, statistics are reported in Fig. 8. As a function of the NPs size, oxygen coordination numbers of Eu and Er are discriminated according to their occurrence and their oxygen species composition (Fig. 8a). As it has been observed for Mg, these data clearly show that oxygen coordination of RE increases with the NPs size in which they are embedded. Within NPs, RE ions encountered the NBO-rich environment induced by the depolymerization of silica, and become mainly surrounded by NBO^{**}. The largest NPs being the richest in Mg and NBO^{**}, RE embedded in large NPs are coordinated with a higher number of oxygen atoms and especially of NBO^{**}. Discriminating the RE sites according to their oxygen coordination number, Fig. 8b shows that the increase of coordination number with the NP size is essentially due to an increase of the proportion of 7- and 8-coordinated sites to the detriment of the 5- and 6-coordinated sites. One can also notice that, whatever the NP size in which RE ions are embedded, 6-coordinated sites are more frequent for Er than for Eu. Likewise, it has been previously shown that non-isolated Mg atoms were mostly coordinated by 6 oxygen atoms. It is also known that a Mg²⁺ ion occupying a 6-coordinated oxygen site can be substituted more easily by an Er³⁺ ion than by a Eu³⁺ ion [28]. This observation is supported by considerations on ionic radii [29, 30]: Er³⁺ ionic radius, ranging from 0.89 Å (VI) to 0.945 Å (VII), is shorter than the one of Eu³⁺ (0.947 Å (VI) – 1.01 Å (VII)), and consequently closer to the one of Mg²⁺ (0.72 Å). Reminding that the largest NPs are composed of 3.3% of ErO_{3/2} against 2.3% of EuO_{3/2} for the same nominal concentration, a stronger affinity of Er³⁺ ions for 6-coordinated sites, which are prevalent in most of NPs, may explain the better encapsulation of erbium compared with europium. In addition, these statistics are in good agreement with the “crystal-chemistry” model proposed by Wang *et al.* [31], where the ratio $r_{RE^{3+}}/r_{O^{2-}} < 0.732$ indicates that RE should be six-coordinated by oxygen (or seven-coordinated at very high rare-earth concentration) and the smaller RE ionic radius provides higher sites regularity.

[Fig 8 around here]

3.4. Rare-earth clustering

Because RE are over-concentrated in NPs, the question of the clustering has to be addressed. Indeed, it is well known that RE clustering occurs in silica matrix for concentrations beyond some hundreds of ppm [32, 33]. Thus, we can here focus on the RE clustering within NPs. To study the degree of RE clustering, RE ions are categorized according to the size of the cluster they form: two RE ions are considered to belong to a

same cluster if they are separated by less than 5.0 Å, which is the distance corresponding to the position of the first minimum of the Er-Er and Eu-Eu RDF curves. In such a way, isolated RE ions are called 1 RE, RE ions forming a pair are denoted 2 RE, and 3+ RE indicates RE ions within clusters composed of three or more RE ions. Statistics are reported in Table 5 for the two codoped glass models and compared with doped glass models (without Mg). At first glance, the Mg-codoping seems to induce an increase of the RE clustering trend. However, one has to keep in mind that Mg-codoping induces Mg-rich phases, i.e. the so-called NPs, where local RE concentration is much more than the nominal concentration of 1 %. Therefore, since the RE clustering is known to be favored by an increase of the RE concentration [25], it is expected to observe a global increase of the RE clustering when the RE are locally found over-concentrated in NPs. But the addressed question is: “Does the clustering degree remain in proportion with the RE concentration in the NPs?”.

[Tab 5 around here]

[Tab 9 around here]

To answer this question, statistics on the RE clustering as a function of the NPs size have been computed and are reported in Fig. 9. For the sake of comparison, doped glass structures without Mg have been simulated with a RE concentration equivalent to the one found in the largest NPs of the codoped structures. Firstly, we observe that the fraction of isolated RE ions tends to decrease with the increase of the NPs size. Consequently, a larger fraction of RE ions are found to form cluster in the largest NPs. This is in agreement with the fact that RE concentration is more important in the largest NP than in the smaller ones. For instance, Er concentration reaches 3.3% into the largest NPs (composed of 126 Mg and more) wherein Er ions are 51.7% isolated and 22.1% form clusters of three and more RE. Comparing these data with those obtained from glass models of 0.023EuO_{3/2}-0.977SiO₂ and 0.033ErO_{3/2}-0.967SiO₂ compositions, one can observe that the clustering trend is significantly restrained. Indeed, for a same ErO_{3/2} concentration of 3.3%, 51.7% of the Er ions remain isolated in the largest NPs whereas only 44.8% are isolated in the structure without Mg. Likewise, clusters of 3 or more RE ions are less prevalent in the large NPs than in the structure without Mg. Moreover, the same trend is observed for Eu and Er. In order to visualize this effect, Fig. 10 displays snapshots of the 0.033ErO_{3/2}-0.967SiO₂ and 0.01ErO_{3/2}-0.10MgO-0.89SiO₂ glass models where only Er ions are represented with a color code depending on the size of the cluster they form. For the sake of comparison, only Er ions involved in the largest NPs are represented in the codoped structure. From these

snapshots and data of Fig. 9, we can assert that the presence of Mg-silicate NPs significantly prevents the formation of large RE clusters, which can be interpreted as a limiting clustering effect due to the oxygen enrichment of the local environment of RE ions embedded in the NPs.

4. Conclusions

Using MD simulations of RE-doped and RE-Mg-codoped silica glasses, it was found that a phase separation mechanism leads to the formation of Mg-rich phases in which the RE ions are mostly embedded. Such heterogeneities, also observed experimentally by APT, can be considered as amorphous NPs of various sizes. This study has pointed out that the chemical composition of these NPs depends on their size, with Mg and RE concentrations increasing with the NP size. A detailed analysis of the NP structure and composition has revealed that the largest NP are over-concentrated in RE while the expected clustering trend is however limited. It was also observed that Er ions are encapsulated in a larger amount than Eu ions. The oxygen-rich environment found in the largest NPs is mostly due to the presence of non-bridging oxygen atoms brought by the Mg-silicate phase. It explains RE over-concentration in the largest NPs since in such NBO-rich environments the luminescent ions can satisfy higher oxygen coordination numbers. Due to ionic radius considerations, Er ions can substitute Mg ions more easily than Eu ions, explaining therefore a better encapsulation of Er ions in the NPs. Moreover, the clustering trend is limited because RE ions have less need to aggregate themselves to share NBO in the Mg-silicate phase.

These results fully contribute to a better understanding of the size-dependent composition of NPs formed *in situ* in a silica matrix. With the aim of tailoring spectroscopic properties of luminescent ions in nanocomposite glasses, the ability to control their close environment through the size and composition of nanoparticles can make this study a relevant guide for the development of future materials. Works are in progress to deduce spectroscopic characteristics, through crystal field parameters calculation, from the different sites occupied by the luminescent ions in these nanostructured glass models.

Declarations of interest

None.

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Figure Captions

Fig. 1. (a) Snapshots of the 0.10MgO-0.90SiO₂ glass structure obtained by MD (oxygen atoms are not represented for the sake of clarity). (b) Comparison with APT analysis (extract from [9]), where Mg atoms are in magenta and silicon atoms are in blue.

Fig. 2. Mg-Mg RDF and Mg-Mg CDF in the three glass models.

Fig. 3. Snapshots showing: (a) 0.10MgO-0.90SiO₂ glass model, (b) 0.01ErO_{3/2}-0.10MgO-0.89SiO₂ glass model, and (c) 0.01EuO_{3/2}-0.10MgO-0.89SiO₂ glass model. Only Mg atoms are represented with a color code related to the size of the nanoparticle (Si, O and RE atoms are not represented for the sake of clarity).

Fig. 4. Mg-O RDF and CDF as a function of NP size (number of Mg atoms) in the 0.10MgO-0.90SiO₂ glass model.

Fig. 5. Different oxygen species contributions to the Mg-O RDF in the 0.10MgO-0.90SiO₂ glass model, on the left side, and schematic representation of the different oxygen species (red, grey and green spheres represent oxygen, silicon and magnesium atoms, respectively), on the right side.

Fig. 6. RE-Mg and RE-Si RDF and CDF (RE = Eu and Er) in the two codoped-glass models.

Fig. 7. Eu-O and Er-O RDF and CDF with and without Mg-codoping.

Fig. 8. Details of RE-O coordination numbers as a function of the NP size in the 0.01REO_{3/2}-0.10MgO-0.89SiO₂ glass model (Eu/Er-O cut-off = 3.0/2.9 Å): (a) distribution according to oxygen species and (b) distribution according to coordinated sites.

Fig. 9. (a) RE clustering in the 0.01REO_{3/2}-0.10MgO-0.89SiO₂ glass model as a function of the NP size. (b) Comparison between the RE clustering found in the largest NPs (126 Mg and more) and the one occurring in a doped glass with a RE concentration equivalent to the one found in the largest NPs of the codoped structures.

Fig. 10. Snapshots of (a) all Er ions in the 0.033ErO_{3/2}-0.967SiO₂ glass model, and (b) of Er ions belonging to the largest NPs (126 Mg and more) in the 0.01ErO_{3/2}-0.10MgO-0.89SiO₂ glass model. RE are represented with a color code depending on the size of the cluster they form. Cyan segments represent Si-O bonds and magenta segments represent Mg-O bonds in the NPs.

Tables

Table 1

Potential parameters for MD modelling of yREO_{3/2}-xMgO-(1-x-y)SiO₂ structures. Values given in exponent of ion type are partial charges. q and q': each oxygen partial charge depends on its own cationic environment.

Table 2

Nanoparticles composition as a function of their size.

Table 3

Details of Qⁿ distributions of Si atoms within the NPs of the 0.10MgO-0.90SiO₂ glass model.

Table 4

Mg-O nearest neighbor distances r, Mg-O coordination numbers N and their distribution in 0.10MgO-0.90SiO₂ glass model as a function of the NP size (Mg-O cut-off = 2.9 Å).

Table 5

Global statistics on the RE clustering for the codoped and doped glass models.

Table 1.

	D_{ij} (eV)	a_{ij} (\AA^{-1})	r_0 (\AA)	C_{ij} (eV. \AA^{12})
Si ^{2.4} -Si ^{2.4}	-	-	-	-
Si ^{2.4} -O ^{-1.2}	0.340554	2.006700	2.100000	1.0
Mg ^{1.33} -Mg ^{1.33}	-	-	-	-
Mg ^{1.33} -Si ^{2.4}	-	-	-	-
Mg ^{1.33} -O ^{-1.33}	0.209290	1.376871	2.733041	1.0
RE ^{2.00} -RE ^{2.00}	-	-	-	-
RE ^{2.00} -Si ^{2.4}	-	-	-	-
Eu ^{2.00} -O ^{-1.33}	0.000373	2.478247	3.770304	1.0
Er ^{2.00} -O ^{-1.33}	0.000811	2.880949	3.344240	1.0
O ^q -O ^q	0.042395	1.379316	3.618701	22.0

Table 2.

NP size (number of Mg atoms)	0.10MgO-0.90SiO ₂	0.01REO _{3/2} -0.10MgO-0.89SiO ₂ RE = Eu ³⁺ / Er ³⁺		
	% [MgO]	% [SiO ₂]	% [MgO]	% [REO _{3/2}]
2-27	17.2	80.3 / 80.6	17.1 / 16.9	2.6 / 2.5
28-65	25.7	72.4 / 72.5	25.3 / 25.0	2.3 / 2.5
66-125	28.6	69.9 / 69.0	27.9 / 28.2	2.2 / 2.8
126+	31.0	68.6 / 66.9	29.1 / 29.8	2.3 / 3.3

Table 3.

NP size (number of Mg atoms)	Q ⁿ [%]				
	Q ⁰	Q ¹	Q ²	Q ³	Q ⁴
0	0.00	0.00	0.00	0.01	99.99
1	0.00	0.03	1.02	23.30	75.41
2-27	0.05	0.28	3.25	31.86	64.42
28-65	0.38	1.88	10.23	37.47	49.74
66-125	0.67	3.73	12.59	38.97	43.96
126+	1.05	5.51	14.17	37.51	41.60

Table 4.

NP size (number of Mg atoms)	r (Å)	N _{Mg-}					N-coordinated sites [%]				
		O	NNO	NBO*	NBO**	BO*	4	5	6	7	8
1	1.92	5.59	0.00	1.89	0.01	3.69	8.81	37.77	37.18	14.48	1.17
2-27	2.01	5.82	0.05	0.79	2.45	2.53	6.01	28.97	44.14	18.26	2.31
28-65	2.03	6.12	0.17	0.40	3.56	1.99	4.32	17.71	43.20	30.89	3.89
66-125	2.04	6.04	0.14	0.35	3.83	1.72	3.68	23.53	41.91	27.02	3.68
126+	2.04	6.19	0.17	0.31	4.03	1.68	2.11	16.43	47.97	27.79	5.53

Table 5.

Glass composition	1 RE [%]	2 RE [%]	3+ RE [%]
0.01EuO _{3/2} -0.10MgO-0.89SiO ₂	65.3	23.2	11.5
0.01EuO _{3/2} -0.99SiO ₂	72.1	20.4	7.5
0.01ErO _{3/2} -0.10MgO-0.89SiO ₂	64.0	21.7	14.3
0.01ErO _{3/2} -0.99SiO ₂	75.8	14.2	10.0

Fig. 1.

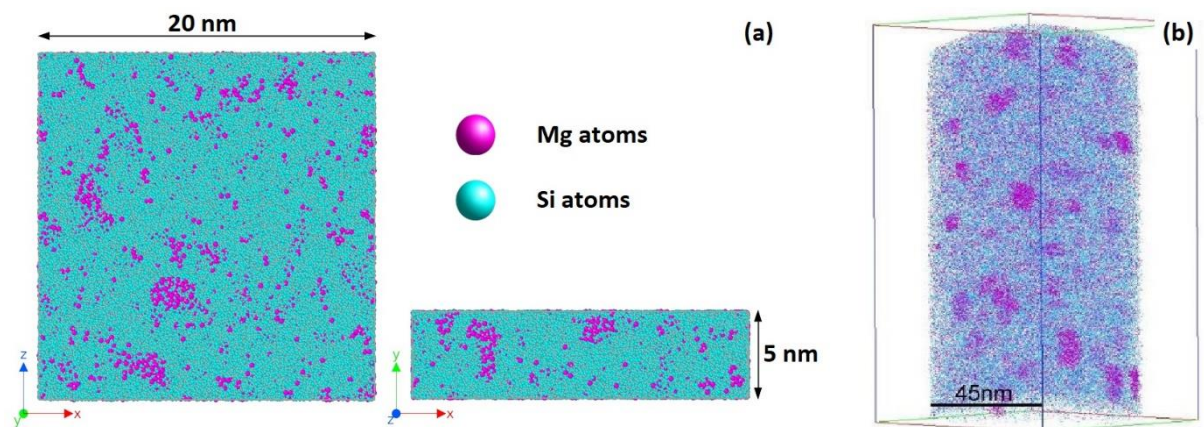


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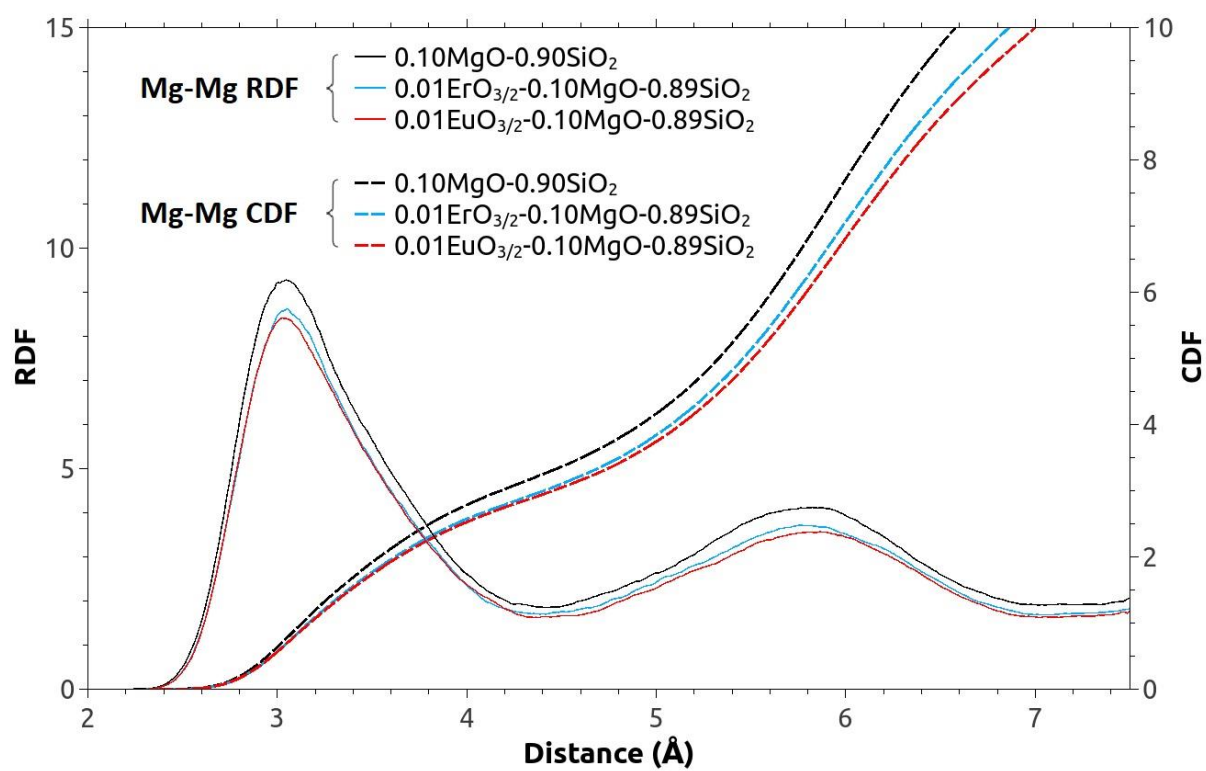
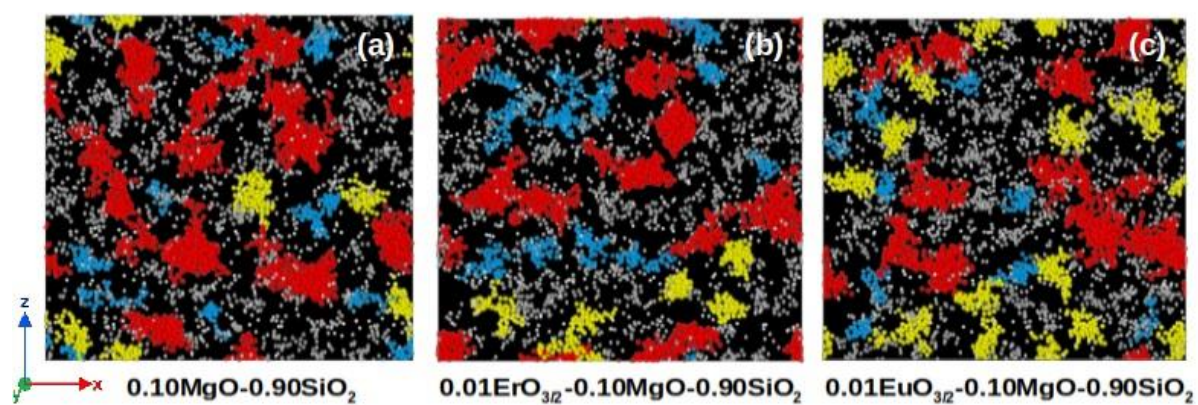


Fig. 3.



NP color code					
NP size (number of Mg atoms)	1	2-27	28-65	66-125	126+

Fig. 4.

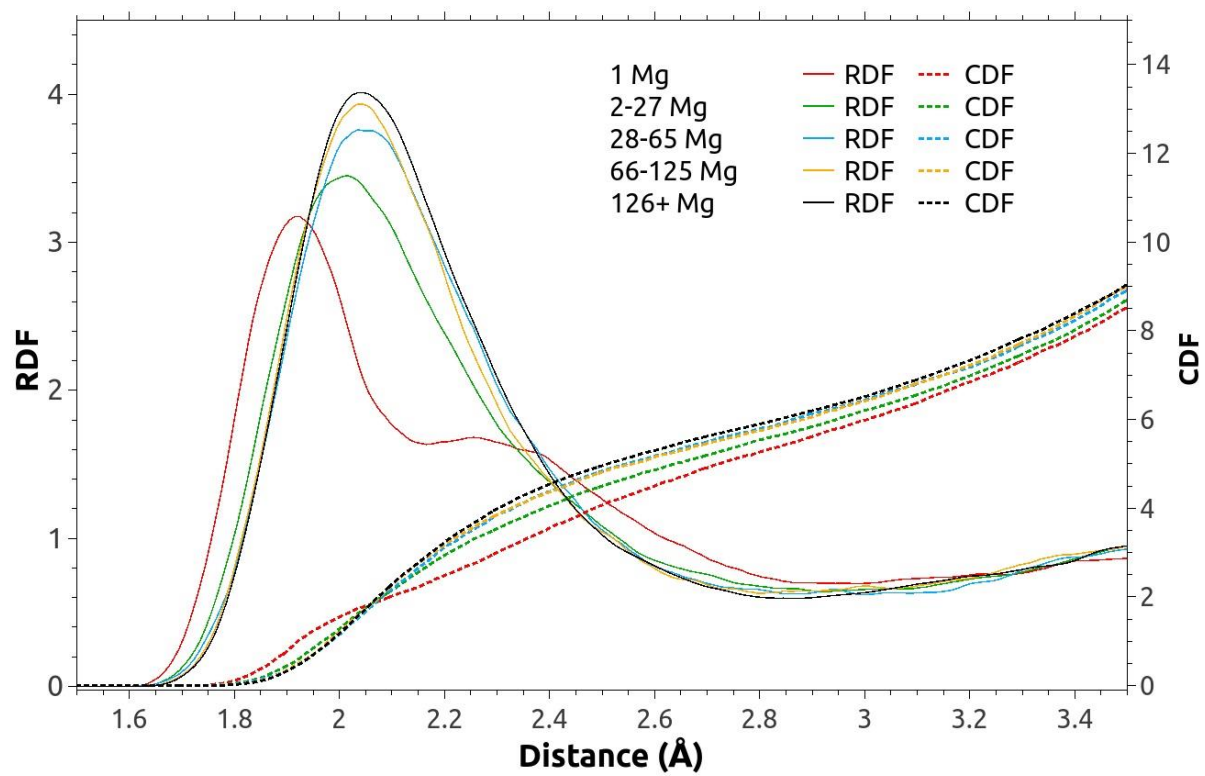


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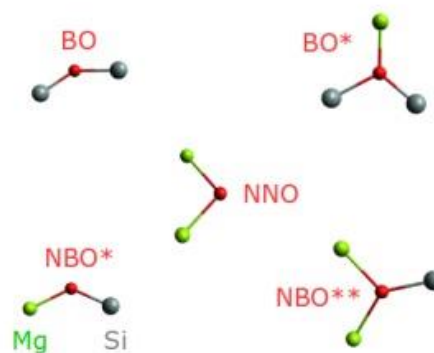
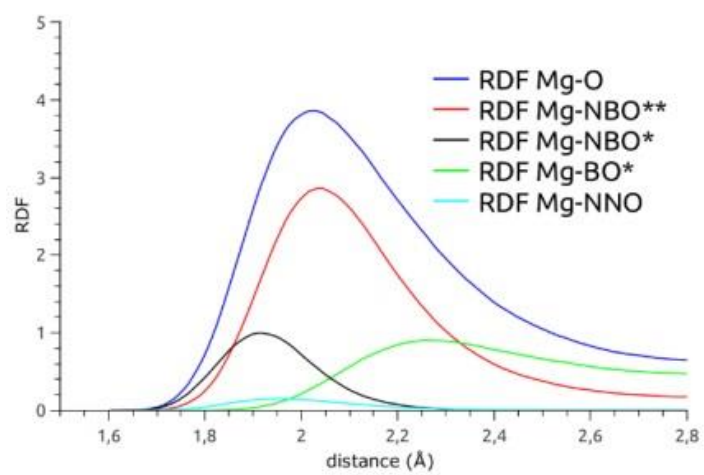


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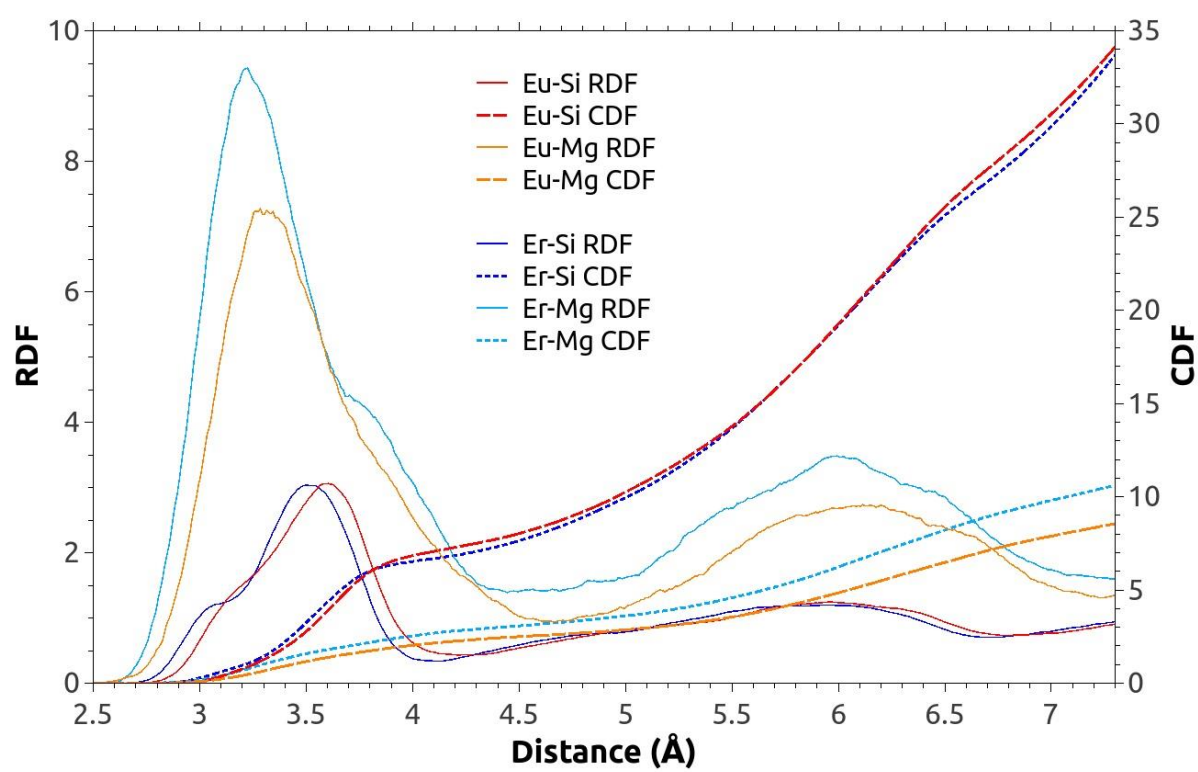


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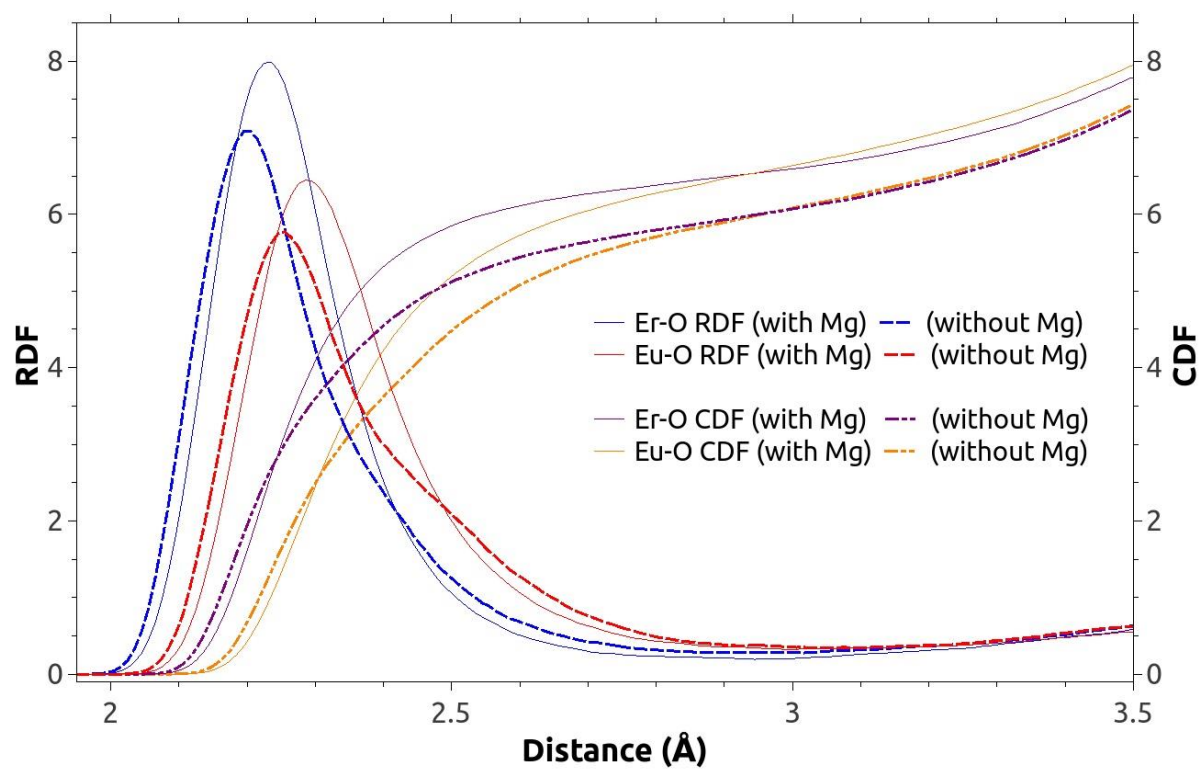


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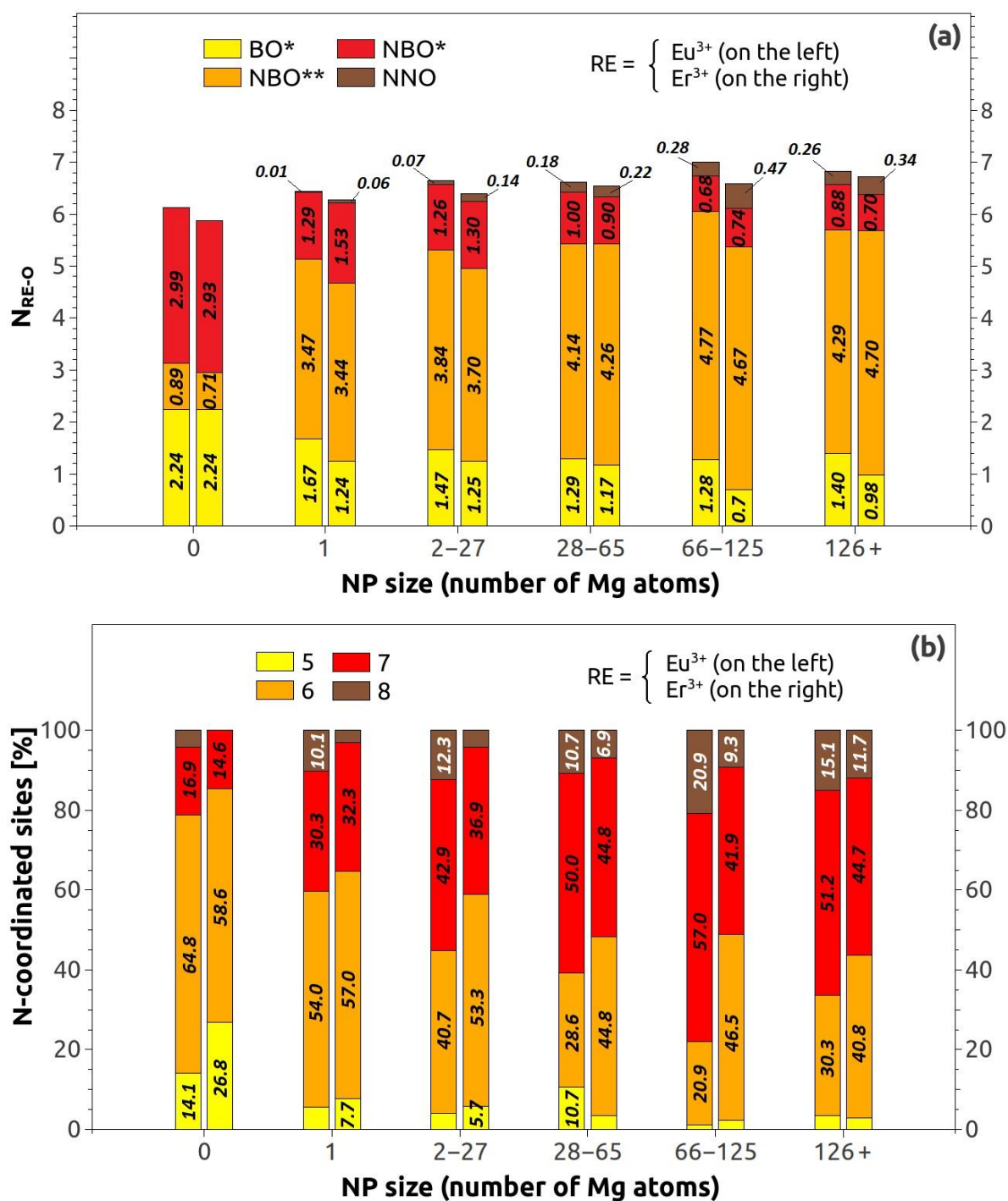


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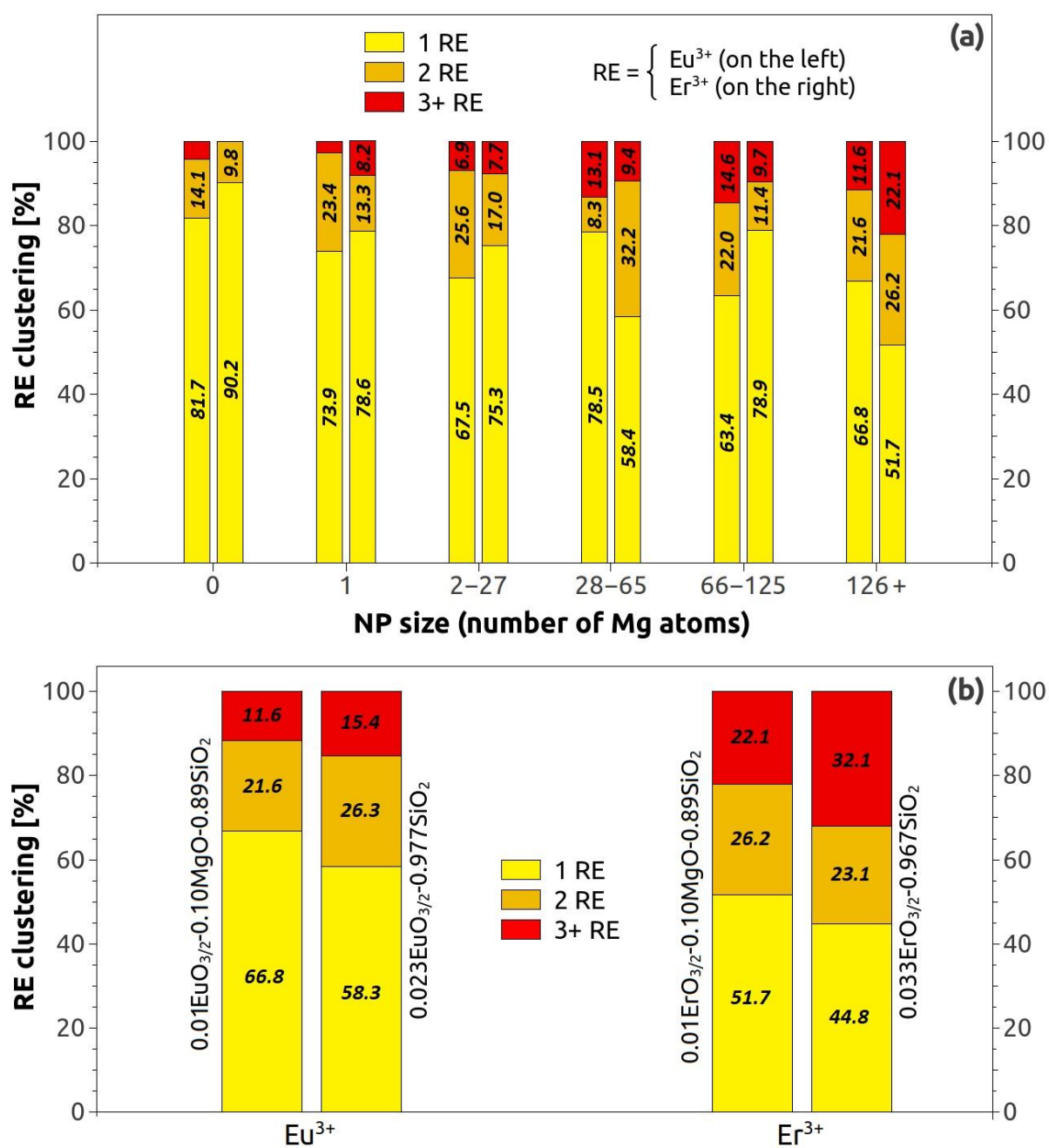


Fig. 10.

